

General Chemistry II Jasperse ANSWERS
 Buffer/Titration/Solubility. Extra Practice Problems

1. C	42. D
2. Answer: A buffer consists of a weak acid and its conjugate base in roughly equal amounts. If acid is added to the solution, it is consumed by the conjugate base. If base is added to the solution, it is consumed by the weak acid. If the amounts are such that the ratio of conjugate base/weak acid concentrations doesn't change much, then the pH doesn't change much. Dilution does not affect the pH because this concentration ratio doesn't change upon dilution.	43. B
3. B	44. D
4. B	45. B
5. C	46. D
6. A	47. B
7. A	48. D
8. D	49. C
9. A	50. B
10. D	51. B
11. B	52. B
12. C	53. D
13. C	54. A
14. B	55. A
15. D	56. B
16. E	57. B
17. D	58. D
18. C	59. B
19. C	60. A
20. 9.43	61. B
21. D	62. D
22. 3.08	63. E
23. 3.61	64. A
24. B	65. B
25. C	66. C
26. A	67. $\text{PbCl}_2(s) \rightarrow \text{Pb}^{2+}(aq) + 2\text{Cl}^{-}(aq)$
27. A	$K_{sp} = [\text{Pb}^{2+}][\text{Cl}^{-}]^2$
28. C	68. B
29. A	69. D
30. D	70. D
31. D	71. B
32. D	72. D
33. A	73. C
34. C	74. C
35. A	75. A
36. 2.74	76. E
37. 3.16	77. D
38. 9.03	78. B
39. 9.55	79. D
40. A	80. C
41. 9.07g	81. A
	82. D
	83. D
	84. C
	85. E
	86. A
	87. C
	88. C
	89. D
	90. E